Isotopes

Isotopes are atoms that have the same number of protons and electrons, but a different number of neutrons. Changing the number of neutrons in an atom does not change the element. Atoms of elements with different numbers of neutrons are called "isotopes" of that element.

Isotopes of Hydrogen are the only ones given specific names.

Protium Deuterium Tritium

All elements have a number of isotopes. Hydrogen has the fewest number of isotopes with only three. The elements with the most isotopes are cesium and xenon with 36 known isotopes.

Isotopes are typically named using their mass.

Carbon-12 is the typical form of the element carbon. However the isotopes Carbon-14 and Carbon-13 also exist. How would these be different?

Carbon-12 Carbon-13 Carbon-14